

**Test Bank for Campbell Biology Canadian 1st Edition by Campbell  
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*Campbell Biology, Cdn. Ed. (Reece et al.)*  
**Chapter 2 The Chemical Context of Life**

1) About 25 of the 92 natural elements are known to be essential to life. Which four of these 25 elements make up approximately 96% of living matter?

- A) carbon, sodium, hydrogen, nitrogen
  - B) carbon, oxygen, phosphorus, hydrogen
  - C) oxygen, hydrogen, calcium, nitrogen
  - D) carbon, hydrogen, nitrogen, oxygen
  - E) carbon, oxygen, nitrogen, calcium
- Answer: D

Type: MC

Topic: Concept 2.1

Skill: Knowledge/Comprehension

2) Trace elements are those required by an organism in only minute quantities. Which of the following is a trace element that is required by humans and other vertebrates, but not by other organisms such as bacteria or plants?

- A) nitrogen
- B) calcium
- C) iodine
- D) sodium
- E) phosphorus

Answer: C

Type: MC

Topic: Concept 2.1

Skill: Knowledge/Comprehension

3) Which of the following statements is *false*?

- A) Carbon, hydrogen, oxygen, and nitrogen are the most abundant elements of living matter.
- B) Some trace elements are very abundant on Earth.
- C) Virtually all organisms require the same elements in the same quantities.

- D) Iron is an example of an element needed by all organisms.  
E) Other than some trace elements, animals are mostly made up of the same elements as plants, in similar proportions.

Answer: C

Type: MC

Topic: Concept 2.1

Skill: Application/Analysis

4) What factors are most important in determining which elements are most common in living matter?

- A) the relative abundances of the elements in Earth's crust and atmosphere  
B) the emergent properties of the simple compounds made from these elements  
C) the reactivity of the elements with water  
D) the chemical stability of the elements  
E) both the relative abundances of the elements and the emergent properties of the compounds made from these elements

Answer: E

Type: MC

Topic: Concept 2.1

Skill: Synthesis/Evaluation

5) Why is each element unique and different from other elements in chemical properties?

- A) Each element has a unique atomic mass.  
B) Each element has a unique atomic weight.  
C) Each element has a unique number of protons in its nucleus.  
D) Each element has a unique number of neutrons in its nucleus.  
E) Each element has different radioactive properties.

Answer: C

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

6) Knowing just the atomic mass of an element allows inferences about which of the following? A) the chemical properties of the element

- B) the number of protons in the element  
C) the number of neutrons in the element  
D) the number of protons plus neutrons in the element  
E) both the number of protons and the chemical properties of the element

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

7) In what way are elements in the same column of the periodic table the same?

- A) They have the same number of protons.  
B) They have the same number of neutrons.  
C) They have the same number of electrons.

D) They have the same number of electrons in their valence shell.

E) They have the same number of electron shells.

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

8) Oxygen has an atomic number of 8 and a mass number of 16. Thus, what is the atomic mass of an oxygen atom?

A) exactly 8 grams

B) exactly 8 daltons

C) approximately 16 grams

D) approximately 16 daltons

E) 24 amu (atomic mass units)

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

9) The nucleus of a nitrogen atom contains 7 neutrons and 7 protons. Which of the following is a *correct* statement concerning nitrogen?

A) The nitrogen atom has a mass number of approximately 7 daltons and an atomic mass of 14.

B) The nitrogen atom has a mass number of approximately 14 daltons and an atomic mass of 7.

C) The nitrogen atom has a mass number of 14 and an atomic mass of 7 grams.

D) The nitrogen atom has a mass number of 7 and an atomic number of 14.

E) The nitrogen atom has a mass number of 14 and an atomic mass of approximately 14 daltons. Answer: E

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

10) Molybdenum has an atomic number of 42. Several common isotopes exist, with mass numbers of 92, 94, 95, 96, 97, 98, and 100. Therefore, which of the following can be *true*? A) Molybdenum atoms can have between 50 and 58 neutrons.

B) The isotopes of molybdenum have different electron configurations.

C) The isotopes of molybdenum can have between 50 and 58 protons.

D) The isotopes of molybdenum have between 50 and 58 neutrons and have different electron configurations.

E) The isotopes of molybdenum have between 50 and 58 protons and have different electron configurations.

Answer: A

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

11) Carbon-12 is the most common isotope of carbon, and has an atomic mass of 12 daltons. A mole of carbon in naturally occurring coal, however, weighs slightly more than 12 grams.

Why? A) The atomic mass does not include the mass of electrons.

B) Some carbon atoms in nature have an extra proton.

C) Some carbon atoms in nature have more neutrons.

D) Some carbon atoms in nature have a different valence electron distribution. E)

Some carbon atoms in nature have undergone radioactive decay.

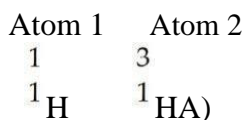
Answer: C

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

12) Which of the following best describes the relationship between the atoms described below?



D) They contain 1 and 3 protons, respectively.

E) They each contain 1 neutron.

Answer: C

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

13) The precise weight of a mole of some pure elements like silicon (Si) can vary slightly from the standard atomic mass, or even from sample to sample. Why?

A) The element may undergo radioactive decay.

B) The element may react with itself and gain or lose subatomic particles.

C) The atoms of the element form chemical bonds with each other, and that changes the weight of the element.

D) The element may have multiple stable isotopes, and the isotopic composition may vary from sample to sample.

E) The amount of energy absorbed by the element affects the mass of its electrons, and thus the atomic mass can vary slightly.

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

12

14

- 14) One difference between carbon-12 ( ${}^6_{12}\text{C}$ ) and carbon-14 ( ${}^6_{14}\text{C}$ ) is that carbon-14 has
- A) two more protons than carbon-12.
  - B) two more electrons than carbon-12.
  - C) two more neutrons than carbon-12.
  - D) two more protons and two more neutrons than carbon-12.
  - E) two more electrons and two more neutrons than carbon-12.

Answer: C

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

- 15) An atom has 6 electrons in its outer shell. How many unpaired electrons does it have? A) 0

B) 2

C) 4

D) 6

E) 2 or 4

Answer: B

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

- 16) The atomic number of nitrogen is 7. Nitrogen-15 is heavier than nitrogen-14 because the atomic nucleus of nitrogen-15 contains how many neutrons?

A) 6 B) 7

C) 8 D) 12

E) 14

Answer: C

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

- 17) Electrons exist only at fixed levels of potential energy. However, if an atom absorbs sufficient energy, a possible result is that

A) an electron may move to an electron shell farther away from the nucleus.

B) an electron may move to an electron shell closer to the nucleus.

C) the atom may become a radioactive isotope.

D) the atom would become a positively charged ion, or cation, and become a radioactive isotope. E) the atom would become a negatively charged ion, or anion.

Answer: A

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

18) The atomic number of neon is 10. Therefore, which of the following is *most* correct about an atom of neon?

- A) It has 8 electrons in its outer electron shell.
- B) It is inert.
- C) It has an atomic mass of 10 daltons.
- D) It has 8 electrons in its outer electron shell and it is inert.
- E) It has 8 electrons in its outer electron shell, it is inert, and it has an atomic mass of 10 daltons.

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

19) From its atomic number of 15, it is possible to predict that the phosphorus atom has

- A) 15 neutrons.
- B) 15 protons.
- C) 15 electrons.
- D) 8 electrons in its outermost electron shell.
- E) 15 protons and 15 electrons.

Answer: E

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

20) Atoms whose outer electron shells contain 8 electrons tend to

- A) form ions in aqueous solutions.
- B) form hydrogen bonds in aqueous solutions.
- C) be stable and chemically nonreactive, or inert.
- D) be gaseous at room temperature.
- E) be both chemically inert and gaseous at room temperature.

Answer: E

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

21) The atomic number of each atom is given to the left of each of the elements below. Which of

the atoms has the same valence as carbon ( ${}^6_6\text{C}$ )?

- A)  ${}^7_7\text{N}$  nitrogen
- B)  ${}^9_9\text{F}$  fluorine
- C)  ${}^{10}_{10}\text{Ne}$  neon
- D)  ${}^{12}_{12}\text{Mg}$  magnesium
- E)  ${}^{14}_{14}\text{Si}$  silicon

Answer: E

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

- 22) Two atoms appear to have the same mass number. These atoms
- A) must have the same atomic number.
  - B) must have the same number of electrons.
  - C) must have the same chemical properties.
  - D) must have the same number of protons + neutrons.
  - E) must have the same atomic number, the same number of protons + neutrons, the same number of electrons, and the same chemical properties.

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Synthesis/Evaluation

- 23) Fluorine has an atomic number of 9 and a mass number of 19. How many electrons are needed to complete the valence shell of a fluorine atom?

A) 1 B) 3

C) 0 D) 7

E) 9

Answer: A

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

- 24) What is the maximum number of electrons in a single  $2p$  orbital of an atom?

A) 1

B) 2

C) 3

D) 4

E) 5

Answer: B

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

- 25) The organic molecules in living organisms have a measurably lower ratio of carbon-13/carbon-12, two stable isotopes of carbon that comprise approximately 1.1% and 98.9% of atmospheric carbon, respectively. What is a reasonable explanation for this phenomenon?
- A) Photosynthesis preferentially uses carbon dioxide molecules with carbon-12, and the lower carbon-13/carbon-12 ratio propagates through the food chain.
  - B) Carbon dioxide molecules with carbon-13 stay in the upper atmosphere and are less available to terrestrial plants and algae.
  - C) Carbon-13 has a different valence electron configuration and is therefore less chemically reactive than carbon-12.
  - D) Oxygen atoms preferentially react with carbon-13, thereby enriching the atmosphere with carbon dioxide molecules containing carbon-13 atoms.
  - E) Carbon dioxide molecules containing carbon-13 are heavier and sink into the ocean depths, making them less available to living organisms.

Answer: A

Type: MC

Topic: Concepts 2.2, 2.4

Skill: Synthesis/Evaluation

- 26) Phosphorus-32, a radioactive isotope of phosphorus-31 (atomic number 15), undergoes a form of radioactive decay whereby a neutron turns into a proton and emits radiation in the form of an electron. What is the product of such radioactive decay of phosphorus-32?

- A) phosphorus-31
- B) a positively charged phosphorus-31 ion
- C) a negatively charged phosphorus-32 ion
- D) sulphur-32 (atomic number 16)
- E) the conversion of the phosphorus-32 atom into pure energy

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

- 27) An atom with atomic number 12 would have what type of chemical behaviour in bonding with other elements?

- A) It would form ions with a +1 charge.
- B) It would form ions with a +2 charge.
- C) It would form ions with a -1 charge.
- D) It would form ions with a -2 charge.
- E) It would form two covalent bonds with other atoms.

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis



28) If a salamander relied on hydrogen bonds to cling to surfaces, what type of surface would cause the most problems for this animal?

- A) a surface coated with a thin film of water
- B) a surface made with carbon and hydrogen atoms covalently bonded together
- C) a surface made with carbon, hydrogen, and oxygen atoms covalently bonded together
- D) a surface made with carbon, hydrogen, nitrogen, and oxygen atoms covalently bonded together
- E) a surface made with silicon and oxygen atoms covalently bonded together

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Synthesis/Evaluation

29) A covalent chemical bond is one in which

- A) electrons are removed from one atom and transferred to another atom so that the two atoms become oppositely charged.
- B) protons and neutrons are shared by two atoms so as to satisfy the requirements of both atoms.
- C) outer-shell electrons of two atoms are shared so as to satisfactorily fill the outer electron shells of both atoms.
- D) outer-shell electrons of one atom are transferred to fill the inner electron shell of another atom.
- E) an electron occupies a hybrid orbital located between the nuclei of two atoms.

Answer: C

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

30) were allowed to react with atoms of hydrogen (atomic number 1), which of the molecules below would be formed?

- A) S—H
- B) H—S—H
- C) H—S—H  
|  
H
- D) H  
| H—  
S—H  
|  
H
- E) H=S=H

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

31) What is the maximum number of covalent bonds an element with atomic number 8 can make with hydrogen?

- A) 1 B) 2
- C) 3 D) 4
- E) 6

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

32) Nitrogen (N) is much more electronegative than hydrogen (H). Which of the following statements is *correct* about the atoms in ammonia (NH<sub>3</sub>)?

- A) Each hydrogen atom has a partial positive charge; the nitrogen atom has a partial negative charge.
- B) The nitrogen atom has a strong positive charge; each hydrogen atom has a strong positive charge.
- C) Each hydrogen atom has a slight negative charge; the nitrogen atom has a strong positive charge.
- D) The nitrogen atom has a slight positive charge; each hydrogen atom has a slight negative charge.
- E) There are covalent bonds between the hydrogen atoms and polar bonds between each hydrogen atom and the nitrogen atom.

Answer: A

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

33) When two atoms are equally electronegative, they will interact to form

- A) hydrogen bonds.
- B) van der Waals interactions.
- C) polar covalent bonds.
- D) nonpolar covalent bonds.
- E) ionic bonds.

Answer: D

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

34) What results from an unequal sharing of electrons between atoms? A) a nonpolar covalent bond

B) a polar covalent bond

C) an ionic bond

D) a hydrogen bond

E) a hydrophobic interaction

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

35) A covalent bond is likely to be polar when

A) one of the atoms sharing electrons is much more electronegative than the other atom.

B) the two atoms sharing electrons are equally electronegative.

C) oxygen is one of the two atoms sharing electrons.

D) one of the atoms has absorbed more energy than the other atom.

E) the two atoms sharing electrons are different elements.

Answer: A

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

36) Which of the following molecules contains the most polar covalent bond?

A) H<sub>2</sub>

B) O<sub>2</sub>

C) CO<sub>2</sub>

D) H<sub>2</sub>O

E) CH<sub>4</sub>

Answer: D

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

37) In comparing covalent bonds and ionic bonds, which of the following would you expect?

A) An atom can form covalent bonds with multiple partner atoms, but only a single ionic bond with a single partner atom.

B) Covalent bonds and ionic bonds occupy opposite ends of a continuous spectrum, from nearly equal to completely unequal sharing of electrons.

C) Both involve electrical attraction between the electrons of one atom and the nucleus of the other atom.

D) Ionic interactions remain when covalent bonds are broken in water. Ionic bonds are much stronger than covalent bonds.

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Synthesis/Evaluation

38) What is the difference between covalent bonds and ionic bonds?

- A) Covalent bonds are formed between atoms to form molecules; ionic bonds are formed between atoms to form compounds.
- B) Covalent bonds involve the sharing of pairs of electrons between atoms; ionic bonds involve the sharing of single electrons between atoms.
- C) Covalent bonds involve the sharing of electrons between atoms; ionic bonds involve the electrical attraction between atoms.
- D) Covalent bonds involve the sharing of electrons between atoms; ionic bonds involve the sharing of protons between atoms.
- E) Covalent bonds involve the transfer of electrons between atoms; ionic bonds involve the sharing of electrons between atoms.

Answer: C

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

39) In ammonium chloride salt ( $\text{NH}_4\text{Cl}$ ) the anion is a single chloride ion,  $\text{Cl}^-$ . What is the cation of  $\text{NH}_4\text{Cl}$ ?

- A) N, with a charge of +1
- B)  $\text{NH}$ , with a charge of +1
- C)  $\text{H}_3$ , with a charge of +1
- D)  $\text{NH}_4$ , with a charge of +1
- E)  $\text{NH}_4$ , with a charge of +4

Answer: D

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

40) The atomic number of chlorine is 17. The atomic number of magnesium is 12. What is the formula for magnesium chloride?

- A)  $\text{MgCl}$  B)
- $\text{MgCl}_2$  C)
- $\text{Mg}_2\text{Cl}$  D)
- $\text{Mg}_2\text{Cl}_2$  E)
- $\text{MgCl}_3$

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

41) How many electron pairs are shared between carbon atoms in a molecule that has the formula  $C_2H_4$ ?

- A) 0
- B) 1
- C) 2
- D) 3
- E) 4

Answer: C

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

42) Which bond or interaction would be difficult to disrupt when compounds are put into water? A) covalent bond

- B) hydrogen bond
- C) van der Waals interaction
- D) ionic bond
- E) either covalent bonds or ionic bonds

Answer: A

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

43) Which of the following explains most specifically the attraction of water molecules to one another?

- A) nonpolar covalent bond
- B) polar covalent bond
- C) ionic bond
- D) hydrogen bond
- E) hydrophobic

interaction Answer: D

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

44) Van der Waals interactions result when

- A) hybrid orbitals overlap.
- B) electrons are not symmetrically distributed in a molecule.
- C) molecules held by ionic bonds react with water.
- D) two polar covalent bonds react.
- E) a hydrogen atom loses an electron.

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

45) What bonding or interaction is most likely to occur among a broad array of molecules of various types (polar, nonpolar, hydrophilic, hydrophobic)?

- A) covalent bonding
- B) polar covalent bonding
- C) ionic bonding
- D) hydrogen bonding
- E) van der Waals interactions

Answer: E

Type: MC

Topic: Concept 2.3

Skill: Synthesis/Evaluation

46) Which of the following is *not* considered to be a weak molecular interaction?

- A) a covalent bond
- B) a van der Waals interaction
- C) an ionic bond in the presence of water
- D) a hydrogen bond
- E) both a hydrogen bond and a covalent bond

Answer: A

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

47) Which of the following would be regarded as compounds?

- A) H<sub>2</sub>O, O<sub>2</sub>, and CH<sub>4</sub>
- B) H<sub>2</sub>O and O<sub>2</sub>
- C) O<sub>2</sub> and CH<sub>4</sub>
- D) CH<sub>4</sub> and O<sub>2</sub>, but not H<sub>2</sub>O
- E) H<sub>2</sub>O and CH<sub>4</sub>, but not O<sub>2</sub>

Answer: E

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

48) What is the maximum number of hydrogen atoms that can be covalently bonded in a molecule containing two carbon atoms?

- A) 2
- B) 3
- C) 4
- D) 6
- E) 8

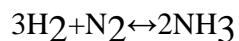
Answer: D

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

49) Which of the following is *true* for this reaction?



- A) The reaction is nonreversible.
- B) Hydrogen and nitrogen are the reactants of the reverse reaction.
- C) Hydrogen and nitrogen are the products of the forward reaction.
- D) Ammonia is being formed and decomposed.
- E) Hydrogen and nitrogen are being decomposed.

Answer: D

Type: MC

Topic: Concept 2.4

Skill: Knowledge/Comprehension

50) Which of the following *correctly* describes chemical equilibrium?

- A) Forward and reverse reactions continue with no effect on the concentrations of the reactants and products.
- B) Concentrations of products are higher than the concentrations of the reactants.
- C) Forward and reverse reactions have stopped so that the concentration of the reactants equals the concentration of the products.
- D) Reactions stop only when all reactants have been converted to products.
- E) There are equal concentrations of reactants and products, and the reactions have stopped.

Answer: A

Type: MC

Topic: Concept 2.4

Skill: Knowledge/Comprehension

51) Which of the following *correctly* describes any reaction that has reached chemical equilibrium? A) The concentration of the reactants equals the concentration of the products.  
B) The rate of the forward reaction is equal to the rate of the reverse reaction.  
C) All of the reactants have been converted to the products of the reaction.  
D) All of the products have been converted to the reactants of the reaction.  
E) Both the forward and the reverse reactions have stopped with no net effect on the concentration of the reactants and the products.

Answer: B

Type: MC

Topic: Concept 2.4

Skill: Knowledge/Comprehension

52) Which of these systems is least likely to be at chemical equilibrium? A) a test tube of living cells  
B) a test tube of organic molecules, kept in the freezer  
C) a test tube of dry organic molecules, kept at room temperature  
D) a test tube of organic molecules dissolved in water, kept at room temperature  
E) a test tube of dead cells in water, kept at room temperature

Answer: A

Type: MC

Topic: Concept 2.4

Skill: Synthesis/Evaluation

53) The combining of the metal, sodium, with the poisonous gas, chlorine, to produce an edible product, salt, is a good example of

- A) essential elements.
- B) emergent properties.
- C) covalent interactions.
- D) Van der Waals interactions.
- E) chemical equilibrium.

Answer: B

Type: MC

Topic: Concept 2.1

Skill: Application/Analysis

54) Plants that are capable of thriving in serpentine soil can do so as a result of

- A) natural selection.
- B) chance.
- C) chemical neutralization of contaminants.
- D) generating their own essential elements.
- E) serpentine soil poses no challenge to plants.

Answer: A

Type: MC

Topic: Concept 2.1

Skill: Knowledge/Comprehension

55) The three types of subatomic particles pertinent to the study of biology are

- A) electrons, protons, and neutrinos.
- B) electrons, positrons, and neutrons.
- C) electrons, protons, and neutrons.
- D) electrons, photons, and neutrons.
- E) quarks, photons, and gravitons.

Answer: C

Type: MC

Topic: Concept 2.1

Skill: Knowledge/Comprehension



56) A dalton is a unit of

A) distance.

B) weight.

C) mass.

D) bond

strength. E)

energy. Answer:

C Type: MC

Topic: Concept 2.1

Skill: Knowledge/Comprehension

57) Chemical bond/interaction strength appears in what order? A) Hydrogen > covalent > ionic > VanderWaals.

B) Covalent > ionic > hydrogen > VanderWaals.

C) Covalent > hydrogen > ionic > VanderWaals.

D) VanderWaals > hydrogen > ionic > covalent.

E) Ionic > hydrogen > VanderWaals > covalent.

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

First shell	Hydrogen ${}^1_1\text{H}$ 	<div style="display: flex; align-items: center; justify-content: center;"> <div style="border: 1px solid black; padding: 2px; margin-right: 10px;"> <math>{}^2_2\text{He}</math>            He            4.00         </div> <div style="margin-right: 10px;">             Atomic number              Element symbol              Mass number           </div> <div style="margin-right: 10px;">             Helium  <math>{}^2_2\text{He}</math>               Electron distribution diagram           </div> </div>						Helium ${}^2_2\text{He}$ 
Second shell	Lithium ${}^3_3\text{Li}$ 	Beryllium ${}^4_4\text{Be}$ 	Boron ${}^5_5\text{B}$ 	Carbon ${}^6_6\text{C}$ 	Nitrogen ${}^7_7\text{N}$ 	Oxygen ${}^8_8\text{O}$ 	Fluorine ${}^9_9\text{F}$ 	Neon ${}^{10}_{10}\text{Ne}$ 
Third shell	Sodium ${}^{11}_{11}\text{Na}$ 	Magnesium ${}^{12}_{12}\text{Mg}$ 	Aluminium ${}^{13}_{13}\text{Al}$ 	Silicon ${}^{14}_{14}\text{Si}$ 	Phosphorus ${}^{15}_{15}\text{P}$ 	Sulfur ${}^{16}_{16}\text{S}$ 	Chlorine ${}^{17}_{17}\text{Cl}$ 	Argon ${}^{18}_{18}\text{Ar}$ 

58) Refer to the figure above (first three rows of the periodic table). If life arose on a planet where carbon were absent, which element might fill the role of carbon?

- A) boron
- B) silicon
- C) nitrogen
- D) aluminum
- E) phosphorus

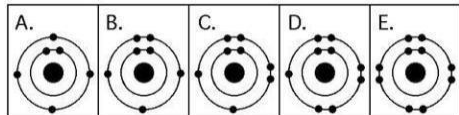
Answer: B

Type: MC

Topic: Concept 2.2

Skill: Synthesis/Evaluation

Use the following figure to answer the questions below.



59) Which drawing in the figure above depicts the electron configuration of an element with chemical properties most similar to Helium ( ${}^2\text{He}$ )?

- A) A
- B) B
- C) C
- D) D
- E) E

Answer: E

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

60) Which drawing in the figure above depicts the electron configuration of an atom that can form covalent bonds with two hydrogen atoms?

- A) A B) B
- C) C D) D
- E) E

Answer: C

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

61) Which drawing in the figure above depicts the electron configuration of an atom capable of forming three covalent bonds with other atoms?

- A) A B) B
- C) C D) D
- E) E

Answer: B

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

Topic: Concept 2.2

Skill: Knowledge/Comprehension

62) Which drawing in the figure above is of the electron configuration of a sodium  $11\text{Na}^+$  ion?

- A) A
- B) B
- C) C
- D) D
- E) E

Answer: E

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

63) Which drawing in the figure above depicts the most electronegative atom? A) A

- B) B C) C
- D) D E) E

Answer: D

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

64) Which drawing in the figure above depicts an atom with a valence of 3? A) A

- B) B C) C
- D) D E) E

Answer: B

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

65) Which drawing in the figure above depicts an atom with a valence of 2? A) A

- B) B C) C
- D) D E) E

Answer: C

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

Topic: Concept 2.2

Skill: Knowledge/Comprehension

Use the following figure to answer the questions below.

Atomic mass →	12	16	1	14	32	31
	C	O	H	N	S	P
Atomic number →	6	8	1	7	16	15

66) In the figure above, how many electrons does nitrogen have in its valence shell? A) 2

B) 5 C) 7

D) 8 E) 14

Answer: B

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

67) In the figure above, how many unpaired electrons does phosphorus have in its valence shell?

A) 15

B) 2 C) 3

D) 7 E) 5

Answer: C

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

68) How many neutrons are present in the nucleus of a phosphorus-32 ( $^{32}_{15}\text{P}$ ) atom (see the figure above)?

A) 5

B) 15

C) 16

D) 17

E) 32

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

69) How many electrons does an atom of sulphur have in its valence shell (see the figure above)?

A) 4

B) 6 C) 8

D) 16 E)

32

Answer: B

Type: MC

Topic: Concept 2.2

Skill: Knowledge/Comprehension

70) Based on electron configuration, which of these elements in the figure above would exhibit a chemical behaviour most like that of oxygen?

A) carbon B)

hydrogen C)

nitrogen D)

sulphur E)

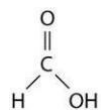
phosphorus

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis



71) The illustration above shows a representation of formic acid. A formic acid molecule

A) will form hydrogen bonds with water molecules.

B) has a tetrahedral configuration of hybrid electron orbitals for the carbon atom.

C) consists of largely nonpolar covalent bonds.

D) is held together by hydrogen bonds.

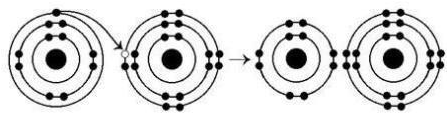
E) has a tetrahedral shape and will form hydrogen bonds with water molecules. Answer: A

Type: MC

Topic: Concept 2.3

Skill: Synthesis/Evaluation

Use the following figure to answer the questions below.



- 72) What results from the chemical reaction illustrated above? A) a cation with a net charge of +1  
B) a cation with a net charge of -1  
C) an anion with a net charge of +1  
D) an anion with a net charge of -1  
E) a cation with a net charge of +1 and an anion with a net charge of -1

Answer: E

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension

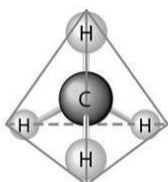
- 73) What is the atomic number of the cation formed in the reaction illustrated above? A) 1  
B) 8 C) 10  
D) 11 E) 16

Answer: D

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis



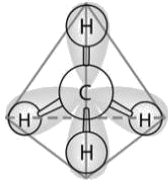
- 74) What causes the shape of the molecule shown above?  
A) the configuration of the 2 *p* orbitals in the carbon atom  
B) the configuration of the 1 *s* orbital in the carbon atom  
C) the configuration of the *sp* hybrid orbitals of the electrons shared between the carbon and hydrogen atoms  
D) the packing of the carbon and hydrogen atoms in a crystal lattice  
E) hydrogen bonding configurations between the carbon and hydrogen atoms

Answer: C

Type: MC

Topic: Concept 2.3

Skill: Knowledge/Comprehension



75) In the methane molecule shown in the figure above, bonds have formed that include both the  $s$  orbital valence electrons of the hydrogen atoms and the  $p$  orbital valence electrons of the carbon. The electron orbitals in these bonds are said to be

- A) double orbitals.
- B) tetrahedral orbitals.
- C) complex orbitals.
- D) hybrid orbitals.
- E) polar orbitals.

Answer: D Type:

MC

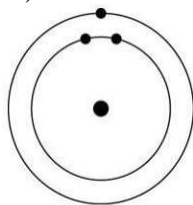
Topic: Concept 2.3

Skill: Knowledge/Comprehension

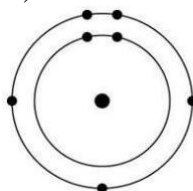


76) Which one of the atoms shown would be most likely to form a cation with a charge of +1?

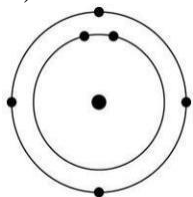
A)



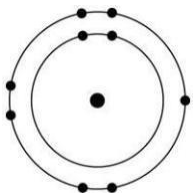
B)



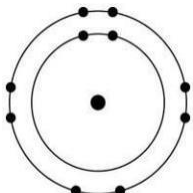
C)



D)



E)



Answer: A

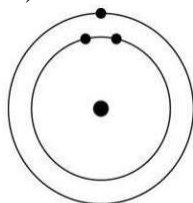
Type: MC

Topic: Concept 2.3

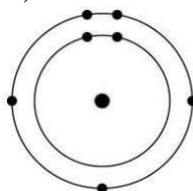
Skill: Application/Analysis

77) Which one of the atoms shown would be most likely to form an anion with a charge of -1?

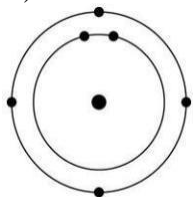
A)



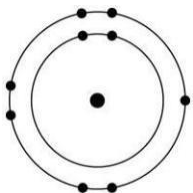
B)



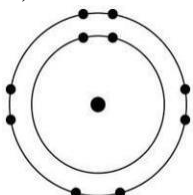
C)



D)



E)



Answer: D

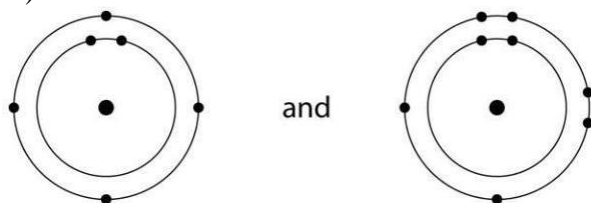
Type: MC

Topic: Concept 2.3

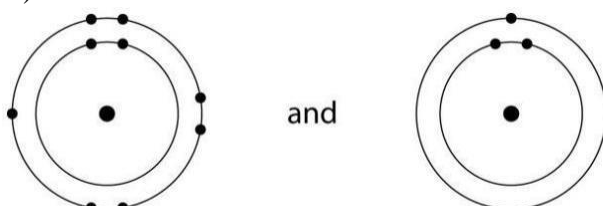
Skill: Application/Analysis

78) Which of the following pairs of atoms would be most likely to form a polar covalent bond?

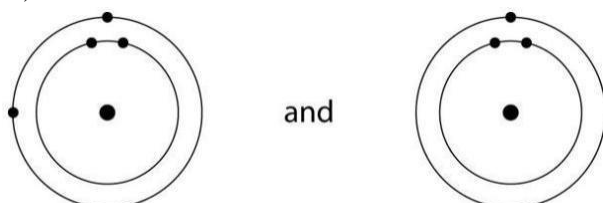
A)



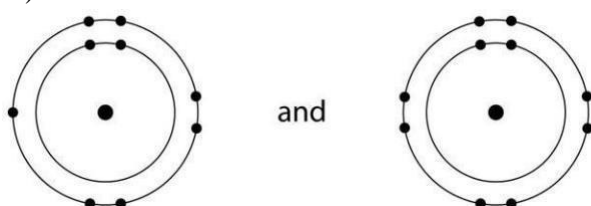
B)



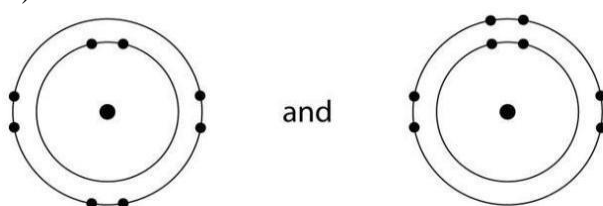
C)



D)



E)



Answer: A

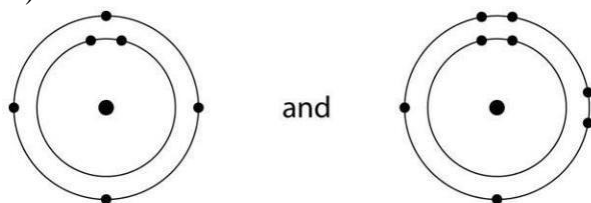
Type: MC

Topic: Concept 2.3

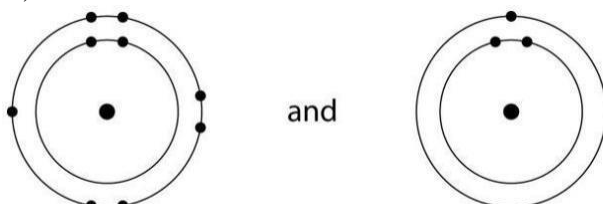
Skill: Application/Analysis

79) Which of the following pairs of atoms would be most likely to form an ionic bond?

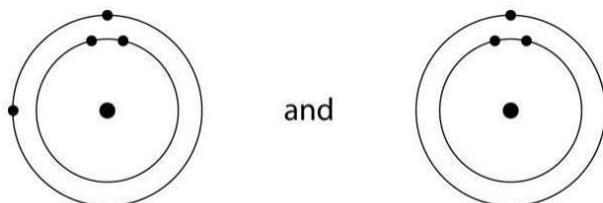
A)



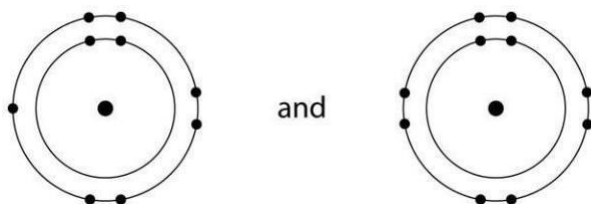
B)



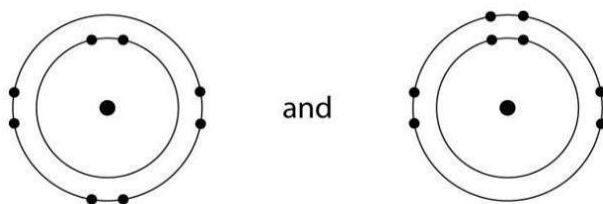
C)



D)



E)



Answer: B

Type: MC

Topic: Concept 2.3

Skill: Application/Analysis

80) A group of molecular biologists is trying to synthesize a new artificial compound to mimic the effects of a known hormone that influences sexual behaviour. They have turned to you for advice. Which of the following compounds is most likely to mimic the effects of the hormone?

- A) a compound with the same number of carbon atoms as the hormone
- B) a compound with the same molecular mass (measured in daltons) as the hormone
- C) a compound with the same three-dimensional shape as part of the hormone
- D) a compound with the same number of orbital electrons as the hormone
- E) a compound with the same number of hydrogen and nitrogen atoms as the hormone

Answer: C

Type: MC

Topic: Concept 2.4

Skill: Application/Analysis

Use the following information to answer the questions below.

You are investigating how chemical reactions occur. You place two reactants together and measure the concentration of product at regular intervals. After a time, the amount of product becomes stable.

81) This solution has

- A) used up all the reactants, so no more product can be made.
- B) used up all the product, so no more reaction is occurring.
- C) reached equilibrium, where there is no more formation of the product.
- D) reached equilibrium, where the net formation of both product and reactants is neutral.
- E) become saturated.

Answer: D

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

82) If you add more product to the solution, you would expect to see

- A) a precipitation of the product.
- B) an increase in pH.
- C) the reactant concentration to remain the same.
- D) the reactant concentration to decrease.
- E) the reactant concentration to increase.

Answer: E

Type: MC

Topic: Concept 2.2

Skill: Application/Analysis

83) In the term *trace element*, the modifier *trace* means that  
A) the element is required in very small amounts.  
B) the element can be used as a label to trace atoms through an organism's metabolism. C) the element is very rare on Earth.  
D) the element enhances health but is not essential for the organism's long-term survival. E) the element passes rapidly through the organism.

Answer: A

Type: MC

Topic: End-of-Chapter Questions

Skill: Knowledge/Comprehension

84) Compared with  $^{31}\text{P}$ , the radioactive isotope  $^{32}\text{P}$  has

A) a different atomic number.

B) a different charge.

C) one more proton.

D) one more electron.

E) one more neutron.

Answer: E

Type: MC

Topic: End-of-Chapter Questions

Skill: Knowledge/Comprehension

85) The reactivity of an atom arises from  
A) the average distance of the outermost electron shell from the nucleus.  
B) the existence of unpaired electrons in the valence shell.  
C) the sum of the potential energies of all the electron shells.  
D) the potential energy of the valence shell.  
E) the energy difference between the *s* and *p* orbitals.

Answer: B

Type: MC

Topic: End-of-Chapter Questions

Skill: Knowledge/Comprehension

86) Which statement is *true* of all atoms that are anions? A) The atom has more electrons than protons.  
B) The atom has more protons than electrons.  
C) The atom has fewer protons than does a neutral atom of the same element. D) The atom has more neutrons than protons.  
E) The net charge is 1-.

Answer: A

Type: MC

Topic: End-of-Chapter Questions

Skill: Knowledge/Comprehension

87) Which of the following statements *correctly* describes any chemical reaction that has reached equilibrium?

- A) The concentrations of products and reactants are equal.
- B) The reaction is now irreversible.
- C) Both forward and reverse reactions have halted.
- D) The rates of the forward and reverse reactions are equal.
- E) No reactants remain.

Answer: D

Type: MC

Topic: End-of-Chapter Questions

Skill: Knowledge/Comprehension

88) We can represent atoms by listing the number of protons, neutrons, and electrons— for example,

$2p^+ ; 2n^0 ; 2e^-$  for helium. Which of the following represents the  $^{18}\text{O}$  isotope of oxygen?

- A)  $6p^+ , 8n^0 , 6e^-$
- B)  $8p^+ , 10n^0 , 8e^-$
- C)  $9p^+ , 9n^0 , 9e^-$
- D)  $7p^+ , 2n^0 , 9e^-$
- E)  $10p^+ , 8n^0 , 9e^-$

Answer: B Type:

MC

Topic: End-of-Chapter Questions

Skill: Application/Analysis

89) The atomic number of sulphur is 16. Sulphur combines with hydrogen by covalent bonding to form a compound, hydrogen sulfide. Based on the number of valence electrons in a sulphur atom, predict the molecular formula of the compound.

- A) HS
- B) HS<sub>2</sub>
- C) H<sub>2</sub>S
- D) H<sub>3</sub>S<sub>2</sub>
- E) H<sub>4</sub>S

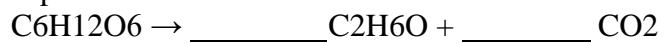
Answer: C

Type: MC

Topic: End-of-Chapter Questions

Skill: Application/Analysis

90) What coefficients must be placed in the following blanks so that all atoms are accounted for in the products?



A) 1; 2

B) 3; 1

C) 1; 3

D) 1; 1

E) 2; 2

Answer: E

Type: MC

Topic: End-of-Chapter Questions

Skill: Application/Analysis